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Japan by the Portuguese, 1543: Copernicus publishes his theory that the Earth and the other planets revolve around the Sun 1545: Theory of complex numbers is first developed by Gerolamo Cardano of Italy. 1558: Camera obscura is first used in Europe by Giambattista della Porta of Italy. 1559-1562: Spanish settlements in Alabama/Florida and Georgia confirm dangers of hurricanes and local native warring tribes. 1565: Spanish settlers outside New Spain (Mexico) colonize Florida's coastline at St. Augustine. 1565: Invention of the graphite pencil (in a wooden holder) by Conrad Gesner. Modernized in 1812. 1568: Gerardus Mercator creates the first Mercator projection map. 1572: Supernova SN 1572 is observed by Tycho Brahe in the Milky Way. 1582: Gregorian calendar is introduced in Europe by Pope Gregory XIII and adopted by Catholic countries. c. 1583: Galileo Galilei of Pisa, Italy identifies the constant swing of a pendulum, leading to development of reliable timekeepers. 1585: earliest known reference to the 'sailing carriage' in China. 1589: William Lee invents the stocking frame. 1591: First flush toilet is introduced by Sir John Harrington of England, the design published under the title 'The Metamorphosis of Ajax'. 1593: Galileo Galilei invents a thermometer. 1596: William Barents discovers Spitsbergen. 1597: Opera in Florence by Jacopo Peri. Entertainment in the 16th century ^ a b Modern reference works on the period tend to follow the introduction of the Gregorian calendar for the sake of clarity; thus NASA's lunar eclipse catalogue states "The Gregorian calendar is used for all dates from 1582 Oct 15 onwards. Before that date, the Julian calendar is used." For dates after 15 October 1582, care must be taken to avoid confusion of the two styles. ^ de Vries, Jan (14 September 2009). "The limits of globalization in the early modern world". *The Economic History Review*. 63 (3): 710–733. CiteSeerX 10.1.1.186.2862. doi:10.1111/j.1468-0289.2009.00497.x. JSTOR 40929823. S2CID 219969360. SSRN 1635517. ^ Singh, Sarina; Lindsay Brown; Paul Clammer; Rodney Cocks; John Mock (2008). *Pakistan & the Karakoram Highway*. Vol. 7, illustrated. Lonely Planet. p. 137. ISBN 978-1-74104-542-0. Retrieved 23 August 2010. ^ Babur (2006). *Babur Nama*. Penguin Books. p. vii. ISBN 978-0-14-400149-1. ^ "16th Century Timeline (1501 to 1600)". fsmitha.com. Archived from the original on February 3, 2009. ^ "History of Smallpox – Smallpox Through the Ages" Archived 2019-09-24 at the Wayback Machine. Texas Department of State Health Services. ^ Ricklefs (1991), p.23 ^ A LIST OF NATIONAL EPIDEMICS OF PLAGUE IN ENGLAND 1348-1665", Archived from the original on 2009-05-08. Retrieved 2009-04-25. ^ a b Ricklefs (1991), page 24 ^ The Sweating Sickness. Story of London. Accessed 2009-04-25. Archived 2009-05-03. ^ Sandra Arlinghaus. "Life Span of Suleiman the Magnificent 1494-1566". Personal.umich.edu. Retrieved 2013-05-05. ^ a b c d e Ricklefs (1991), page 25 ^ "La Terra De Hochelega - Jaques Cartier a Hochelega". jacquescartier.org. Archived from the original on December 23, 2008. ^ "The Lusiads". World Digital Library. 1800-1882. Retrieved 2013-08-31. ^ Schwieger, Peter (2014). *The Dalai Lama and the Emperor of China: a political history of the Tibetan institution of reincarnation*. New York: Columbia University Press. ISBN 9780231538602. OCLC 905914446. ^ Miller, George, ed. (1996). *To The Spice Islands and Beyond: Travels in Eastern Indonesia*. New York: Oxford University Press. pp. xv. ISBN 967-65-3099-9. ^ Luc-Normand Tellier (2009). "Urban world history: an economic and geographical perspective". PUQ. p.308. ISBN 2-7605-1588-5 ^ a b c d e f Ricklefs (1991), page 27 ^ a b Ricklefs (1991), page 28 ^ Polybius: The Rise Of The Roman Empire, Page 36, Penguin, 1979. Langer, William. *An Encyclopedia of World History* (5th ed. 1973); highly detailed outline of events online free Media related to 16th century at Wikimedia Commons Timelines of 16th century events, science, culture and persons Retrieved from " 4 The following pages link to 16th century External tools (link count transclusion count sorted list) · See help page for transcluding these entries Showing 50 items. View (previous 50 | next 50) (20 | 50 | 100 | 250 | 500)Bagpipes (links | edit) List of decades, centuries, and millennia (links | edit) Fashion (links | edit) Giovanni Boccaccio (links | edit) History of Mali (links | edit) History of Mauritius (links | edit) Post office (links | edit) Snare drum (links | edit) Republican Party (United States) (links | edit) 20th century (links | edit) 15th century (links | edit) 17th century (links | edit) 18th century (links | edit) 1624 (links | edit) 1626 (links | edit) 1642 (links | edit) 1661 (links | edit) 1608 (links | edit) 1492 (links | edit) 14th century (links | edit) 1st century (links | edit) 13th century (links | edit) 4th century (links | edit) 12th century (links | edit) 11th century (links | edit) 1564 (links | edit) 1648 (links | edit) 1572 (links | edit) 1623 (links | edit) 1662 (links | edit) 1490s (links | edit) 1640s (links | edit) 1597 (links | edit) 1690 (links | edit) 1688 (links | edit) 1688 (links | edit) 7th century (links | edit) 10th century (links | edit) 9th century (links | edit) 8th century (links | edit) 6th century (links | edit) 5th century (links | edit) 3rd century (links | edit) 2nd century (links | edit) 1573 (links | edit) 1570s (links | edit) 1574 (links | edit) 1436 (links | edit) 1476 (links | edit) 1542 (links | edit) 1540s (links | edit) View (previous 50 | next 50) (20 | 50 | 100 | 250 | 500) Retrieved from " WhatLinksHere/16th century" Exercise 1. An aqueous silver nitrate solution is mixed with an aqueous solution of potassium carbonate. a) Write the chemical formulas for silver nitrate and potassium carbonate AgNO3, K2CO3 b) Write the balanced molecular (overall) equation for the reaction. 2 AgNO3 (aq) + K2CO3 (aq) → Ag2CO3 (s) + 2 KNO3 (aq) c) Write the ionic equation 2 Ag+(aq) + 2 NO3-(aq) + CO32-(aq) → Ag2CO3(s) + 2 K+(aq) + 2 NO3-(aq) d) Write the net ionic equation 2 Ag+(aq) + CO32-(aq) → Ag2CO3(s) e) What are the spectator ions K+ (aq) and NO3-(aq) Back to Precipitation Reactions Exercise 2. An aqueous solution of sodium oxalate is mixed with an aqueous solution of calcium chloride. a) Write the balanced molecular (overall) equation for the reaction. Na2C2O4 (aq) + CaCl2 (aq) → 2 NaCl (aq) + CaC2O4 (s) b) Write the ionic equation 2 Na+ (aq) + C2O4-2 (aq) + Ca2+ (aq) + 2 Cl- (aq)→ CaC2O4 (s) + 2 Na+ (aq) + 2 Cl- (aq) c) Write the net ionic equation Ca2+ (aq) + C2O4-2 (aq) → CaC2O4 (s) d) What are the spectator ions? Na+ and Cl- Back to Precipitation Reactions Exercise 3. An aqueous solution of iron(III) chloride is mixed with an aqueous solution of lead(II) nitrate. a) Write the balanced molecular (overall) equation for the reaction. 2 FeCl3 (aq) + 3 Pb(NO3)2 (aq) → 3 PbCl2 (s) + 2 Fe(NO3)3 (aq) b) Write the ionic equation 2 Fe3+ (aq) + 6 Cl- (aq) + 3 Pb2+ (aq) + 6 NO3- (aq) → 3 PbCl2 (s) + 2 Fe3+ (aq) + 6 NO3- (aq) c) Write the net ionic equation 3 Pb2+ (aq) + 6 Cl- (aq)→ 3 PbCl2 (s) Divide coefficients by 3 and the net ionic equation is: Pb2+ (aq) + 2 Cl- (aq)→ PbCl2 (s) d) What are the spectator ions? Fe3+ and NO3- Back to Precipitation Reactions Exercise 4. Write the net ionic equation for the following. Identify any spectator ions. Fe(ClO4)2 (aq) + (NH4)3PO4)2 (aq) → ?? 3 Fe+2 (aq) + 2 PO43- (aq) → Fe3(PO4)2 (s) The spectator ions are NH4+ and ClO4- Back to Precipitation Reactions Exercise 5. If aqueous solutions of KBr and ZnCrO4 are mixed, will a precipitate form? If so, write the net ionic equation, and identify any spectator ions. The molecular equation is: KBr(aq) + ZnCrO4(aq) → K2CrO4(aq) + ZnBr2 (aq) Notice that both products, K2CrO4 and ZnBr2 are both soluble. There is no reaction-no precipitate is formed. You can determine this from the ionic equation. The ions all cancel. Zn2+ (aq) + CrO42- (aq) + K+ (aq) + Br- (aq) → Zn2+ (aq) + CrO42- (aq) + K+ (aq) + Br- (aq) Back to Precipitation Reactions Exercise 6. A chemist mixes 50.00 mL of aqueous 0.100 M K2SO4 with 45.00 mL of aqueous 0.150 M BaCl2. What is the mass, in g, of the precipitate formed? Step 1. Write the balanced molecular equation. K2SO4 (aq) + BaCl2 (aq) → BaSO4 (s) + 2 KCl (aq) Step 2. Next, determine the moles of K2SO4 and BaCl2 by multiplying the volume, in L, by the molarity. The molar mass of K2SO4 is 174.259 g/mol and the molar mass of BaCl2 is 208.23 g/mol. molesK2SO4= 0.100 M x 0.05000 L= 0.005000 moles K2SO4 molesBaCl2= 0.150 M x 0.04500 L= 0.006750 moles BaCl2 Step 3. Our limiting reactant is K2SO4. We can now use our mole ratios from the balanced chemical equation to determine the moles and then the grams of BaSO4 formed. The molar mass of BaSO4 is 233.38 g/mol. The molar ratio of K2SO4 to BaSO4 is 1:1.

0.005000
mol
K

2

S

O

4

1
mol
;
BaS

O

4

{\displaystyle 0.005000\;mol\;K_{2}SO_{4}\times {\frac {1\;mol\;BaSO_{4}}{1\;mol\;K_{2}SO_{4}}}\;=\;0.005000\;mol\;BaSO_{4}}

 Step 4. We can now calculate the number of grams BaSO4 formed.

0.005000
mol
BaS

O

4

{\displaystyle 0.005000\;mol\;BaSO_{4}}

 Back to Precipitation Reactions Exercise 7. Determine if a precipitation reaction will occur if aqueous solutions of the substances in each pair are mixed. If a precipitation reaction does occur, write a net ionic equation and identify the spectator ions. a) KOH and HClO4 The molecular equation is: KOH (aq) + HClO4(aq) → K+ (aq) + ClO4-(aq) + H2O (l) We see this is the reaction between a strong acid and a strong base that produces a soluble salt and water.No precipitate forms. We have the net ionic equation: H+ (aq) + OH-(aq) → H2O (l) b) Hg(NO3)2 and Na3PO4 The molecular equation is: 3 Hg(NO3)2(aq) + 2 Na3PO4(aq) → Hg3(PO4)2 (s) + 6 NaNO3 (aq) The net ionic equation is: 3 Hg2+ (aq) + 2 PO43- (aq) → Hg3(PO4)2 (s) The spectator ions are Na+ and NO3- c) CoCl2 and Pb(NO3)2 The molecular equation is: CoCl2 + Pb(NO3)2 → PbCl2 (s) + Co(NO3)2 (aq) The net ionic equation is: Pb2+ (aq) + Cl- (aq) → PbCl2 (s) The spectator ions are Co2+ and NO3- d) Sr(OH)2 and BaCl2 The molecular equation is: Sr(OH)2 + BaCl2 → SrCl2 (aq) + Ba(OH)2 (aq) Both SrCl2 and Ba(OH)2 are soluble, therefore NO REACTION Back to Precipitation Reactions Exercise 8. A chemist mixes 150.0 mL of 0.100 M Na2SO4, 75.0 mL 0.350 M CoCl2, and 150.0 mL of 0.250 M SrS. a) Which salts (ionic compounds) will precipitate out of solution? Ions in each solution are Na+ and SO42-, Co2+ and Cl-, Sr2+ and S2-. When the solutions are mixed, CoS and SrSO4 will precipitate. b) Assuming complete precipitation of insoluble salts, what is the concentration, in M, of each ion that remains in solution? 1. Find the number of moles of each ion Moles of Na2SO4 = 0.1500 L x 0.100 M = 0.0150 mol Na2SO4 Moles of Na+ = 2 x 0.0150 mols = 0.0300 mol Na+ Moles of SO42- = 0.0150 mol SO42- Moles of CoCl2 = 0.0750 L x 0.350 M = 0.0263 mol CoCl Moles of Co2+ = 0.0263 mol Co2+ Moles of Cl- = 2 x 0.0263 mols = 0.0526 mol Cl- Moles of SrS = 0.1500 L x 0.250 M = 0.0375 mol SrS Moles of Sr2+ = 0.0375 mol Sr2+ Moles of S2- = 0.0375 mol S2- 2. The precipitation reactions are: 1) Co2+ (aq) + S2- (aq) → CoS (s) 2) Sr2+ + SO4-2 (aq) → SrSO4 (s) For reaction 1) we have 0.0263 mol Co2+ and 0.0375 mol of S2-. The limiting reactant is Co2+. The amount of S2- required for reaction is:

0.0263
mol
Co

2

+
}
times
frac
{
1
mol
;

S

2

}
{
1
mol
;

Co

2

+
}
}
=
0.0263
mol
;

S

2

+
}

 The amount of unreacted Sr2+ is 0.0375 mol - 0.0263 mol = 0.0112 mol S2- For reaction 2 we have 0.0375 mol Sr2+ and 0.0150 mol SO42-. The limiting reactant is SO42-. The amount of Sr2+ required for reaction is

0.0150
mol
SO

4

2

}
times
frac
{
1
mol
;

Sr

2

+
}
{
1
mol
;

SO

4

2

}
}
=
0.0150
mol
;

Sr

2

+
}

 The amount of unreacted Sr2+ is 0.0375 mol - 0.0150 mol = 0.0225 mol Sr2+ 3. Now we can determine the concentration, M, of each ion in solution. All of the SO42- and Co2+ has been reacted, therefore, [SO42-] = [Co2+] = 0. The total volume of solution is: Vtotal = 150.0 mL + 75.0 mL + 150 mL = 375.0 mL = 0.3750 L

0.0150
mol
;

Sr

2

+
}

}

{
0.3750
L
}
=
0.0299
M
;

}

0.0112
mol
;

S

2

}

}

{
0.3750
L
}
=
0.0299
M
;

}

0.03750
L
}
}
=
0.1000
M
;

}

0.0225
mol
}

}

{
0.3750
L
}
=
0.0600
M
;

}

[

Na

+

]
}
=
frac
{
0.0300
mol
}
{
0.3750
L
}
=
0.0800
M
;

}

[

Cl

−

]
}
=
frac
{
0.0526
mol
}
{
0.3750
L
}
=
0.140
M
;

}

 Back to Precipitation Reactions Exercise 9. A chemist has 4 aqueous solutions, each solution containing the following pairs of ions. The solutions were prepared with sodium and potassium salts. Solution 1. Cl- and NO3- Solution 2. Ca2+ and NH4+ Solution 3. Fe2+ and Ba2+ Solution 4. SO42- and CO32- The chemist has the following salts available: Pb(NO3)2, NaOH, CaS, Na2SO4, K2C2O4, and CsNO3. For each solution, indicate the salt that would separate the ions using a precipitation reaction. Write the net ionic equation for each. For solution 1. Adding Pb(NO3)2 would precipitate PbCl2. The chloride would be removed leaving the NO3- ions. The net ionic equation is: Pb2+ (aq) + 2 Cl- (aq) → PbCl2 (s) For solution 2. K2C2O4 would precipitate CaC2O4. The calcium ion would be removed leaving the NH4+ ions. The net ionic equation is: Ca2+ (aq) + C2O42- (aq) → CaC2O4 (s) For solution 3. Na2SO4 would precipitate BaSO4. The barium ion would be removed leaving the Fe2+ ions. The net ionic equation is: Ba2+ (aq) + SO42- (aq) → BaSO4 (s) For solution 4. CaS would precipitate CaCO3. The carbonate ions would be removed leaving the sulfate ions. The net ionic equation is: Ca2+ (aq) + CO32- (aq) → CaCO3 (s) Back to Precipitation Reactions Exercise 10. A chemist mixed 250.00 mL of aqueouod 0.450 M Na3PO4 and with 300.00 mL of aqueous 0.355 M CaCl2? She recovered 10.39 g of a precipitate. What is the percent yield? The balanced molecular equation is: 2 Na3PO4 (aq) + 3 CaCl2 (aq) → Ca3(PO4)2 (s) + 6 NaCl (aq) Next we find the number of moles of CaCl2 and Na3PO4 moles CaCl2 = 0.30000 L x 0.355 M = 0.1065 mol CaCl2 moles Na3PO4 = 0.25000 L x 0.450 M = 0.1125 mol Na3PO4 Our limiting reactant is CaCl2

0.1065
mol
CaCl

2

times
frac
{
1
mol
;

Ca

3

(

P

O

4

)

2

}
{
3
mol
;
CaCl

2

}
}
=
0.0355
mol
;

Ca

3

(

P

O

4

)

2

}

 The molar mass of Ca3(PO4)2 is 310.18 g/mol

0.0355
mol
;

Ca

3

(

P

O

4

)

2

}
}
times
frac
{
310.18
g
;

Ca

3

(

P

O

4

)

2

}
{
1
mol
;

Ca

3

(

P

O

4

)

2

}
}
=
11.0
g
;

Ca

3

(

P

O

4

)

2

}

 The percent yield

percent
;
yield
}
=
frac
{
10.39
g
}
{
11.0
g
}
}
times
;
100
}
=
;
mathbf
{
94.5
%
}

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