

I'm not robot!



2. For Indium-113 (atomic number 49)  
 a. Determine the number of protons, neutrons, and electrons  
 $Z = 49$     $A = 113$     $N = 64$     $e^- = 49$

b. Write its electron configuration using noble gas notation  
 $[Xe] 4f^{14} 5d^{10} 6s^2$

c. How many unpaired electrons does it contain? 0

33. Name the element represented by each of the following electron configurations and determine the number of unpaired electrons each contains:

a. $1s^2 2s^2 2p^3 3s^1 3p^4$	$5s^2 4p^4$	2
b. $[Kr] 5s^1 4d^7$	$4d^7 5s^1$	2
c. $[Xe] 6s^1 4f^{14} 5d^1 6p^1$	$6s^1 5d^1$	2

34. Determine which of the following electron configurations represent an atom in the ground state and identify that element. (Circle choice and name the element)

a.  $1s^2 2s^2 2p^3 3s^1 3p^5$    b.  $1s^2 2s^2 2p^6 3s^2 3p^5$    c.  $1s^2 3s^2 2s^2 2p^3 3p^5$

35. Three isotopes of argon occur in nature, Ar-36, Ar-38, and Ar-40. Calculate the atomic mass of argon to two decimal places, given the following relative atomic masses and abundances of each of the isotopes: argon-36 (35.97 u, 0.337%), argon-38 (37.96 u, 0.063%), and argon-40 (39.96 u, 99.600%).

$$0.00337(35.97) + 0.00063(37.96) + 0.99600(39.96) = 39.95 \text{ u}$$

36. Calculate the mass in grams of each of the following:

a. 3.00 mol Al

$$3.00 \text{ mol Al} \times \frac{26.98 \text{ g}}{1 \text{ mol}} = 80.9 \text{ g}$$

b.  $2.56 \times 10^{24}$  atoms Li

$$2.56 \times 10^{24} \text{ atoms Li} \times \frac{1 \text{ mol}}{6.022 \times 10^{23} \text{ atoms}} \times \frac{6.94 \text{ g}}{1 \text{ mol}} = 29.5 \text{ g Li}$$

In this worksheet, we will practice calculating the relative atomic masses of elements based on the relative abundance of their isotopes. Isotopes are different atoms of the same element that contain the same number of protons and electrons but a different number of neutrons. These are atoms of the same elements but with different mass numbers. Because of this, the mass of an element is given as relative atomic mass (Ar) by using the average mass of all of the isotopes. The relative atomic mass of an element can be calculated by using the percentage abundance values. The percentage abundance of an isotope is either given or can be read off the mass spectrum. Firstly, find the mass of 100 atoms by multiplying the percentage abundance by the mass of each isotope. Secondly, divide by 100 to find the average atomic mass. For example, if you have two isotopes A and B: A sample of oxygen contains the following isotopes: What is the relative atomic mass of oxygen to 2 dp? A 16.00B 17.18C 16.09D 17.00. Answer: The correct answer is A. Total mass of 100 atoms = (99.76 x 16) + (0.04 x 17) + (0.20 x 18) = 1600.44. Mass of 1 atom = 1600.44 ÷ 100 = 16.0044 = 16.00 (2 dp). Here is another example, but this time using a mass spectrum to obtain the information: Calculate the relative atomic mass of boron using its mass spectrum, to 2 dp: Answer: Total mass of 100 atoms = (19.9 x 10) + (80.1 x 11) = 1080.1. Mass of 1 atom = 1080.1 ÷ 100 = 10.801 = 10.80 (2 dp). Page 2 Page 3 Page 4 Thank you for your participation!

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