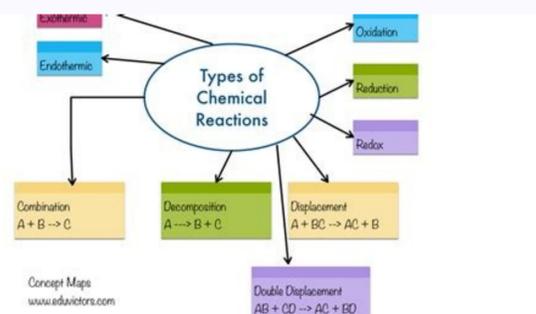


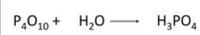
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Balance the following equation:



P = 4 P = 1
O = 11 O = 4
H = 2 H = 3

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Balancing Equations Worksheet Chemistry Project

- $2 NaNO_3 + 1 PbO \rightarrow 1 Pb(NO_3)_2 + 1 Na_2O$
- $6 AgI + 1 Fe_2(CO_3)_3 \rightarrow 2 FeI_3 + 3 Ag_2CO_3$
- $1 C_2H_6 + 2 O_2 \rightarrow 2 CO_2 + 2 H_2O$
- $1 ZnSO_4 + 1 Li_2CO_3 \rightarrow 1 ZnCO_3 + 1 Li_2SO_4$
- $1 V_2O_5 + 5 CaS \rightarrow 5 CaO + 1 V_2S_5$
- $1 Mn(NO_3)_2 + 1 BeCl_2 \rightarrow 1 Be(NO_3)_2 + 1 MnCl_2$
- $3 AgBr + 1 GaPO_4 \rightarrow 1 Ag_3PO_4 + 1 GaBr_3$
- $3 H_2SO_4 + 2 B(OH)_3 \rightarrow B_2(SO_4)_3 + H_2O$
- $1 S_8 + 8 O_2 \rightarrow 8 SO_2$
- $1 Fe + 2 AgNO_3 \rightarrow 1 Fe(NO_3)_2 + 2 Ag$

Chapter 7 Worksheet #1

Balancing Chemical Equations

Balance the equations below:

- $1 N_2 + 3 H_2 \rightarrow 2 NH_3$
- $2 KClO_3 \rightarrow 2 KCl + 3 O_2$
- $2 NaCl + F_2 \rightarrow 2 NaF + Cl_2$
- $2 H_2 + O_2 \rightarrow 2 H_2O$
- $Pb(OH)_2 + 2 HCl \rightarrow 2 H_2O + PbCl_2$
- $2 AlBr_3 + 3 K_2SO_4 \rightarrow 6 KBr + Al_2(SO_4)_3$
- $1 CH_4 + 2 O_2 \rightarrow 1 CO_2 + 2 H_2O$
- $C_3H_8 + 5 O_2 \rightarrow 3 CO_2 + 4 H_2O$
- $C_8H_{18} + 13 O_2 \rightarrow 8 CO_2 + 9 H_2O$
- $FeCl_3 + 3 NaOH \rightarrow 1 Fe(OH)_3 + 3 NaCl$
- $4 P + 5 O_2 \rightarrow P_2O_5$
- $2 Na + 2 H_2O \rightarrow 2 NaOH + H_2$
- $2 Ag_2O \rightarrow 4 Ag + O_2$
- $1 S_8 + 12 O_2 \rightarrow 8 SO_3$
- $CO_2 + H_2O \rightarrow 1 C_6H_{12}O_6 + 6 O_2$
- $1 K + 1 MgBr_2 \rightarrow 1 KBr + 1 Mg$
- $2 HCl + 1 CaCO_3 \rightarrow 1 CaCl_2 + 1 H_2O + 1 CO_2$
- $HNO_3 + NaHCO_3 \rightarrow 1 NaNO_3 + 1 H_2O + 1 CO_2$
- $H_2O + O_2 \rightarrow 2 H_2O_2$
- $NaBr + CaF_2 \rightarrow 2 NaF + 1 CaBr_2$
- $1 H_2SO_4 + 2 NaNO_2 \rightarrow 1 HNO_2 + 1 Na_2SO_4$

Types of Chemical Reactions Worksheet

Writing Formulas
In order to complete chemical reactions, you must be able to write formulas for ionic compounds. Remember, to write the formula for ionic compounds all you have to do is cross the numerical value of the charges into subscripts for the elements. Remember, polyatomic ions must be kept in brackets if there are more than one of them!

- Ex) $Ca^{2+} Cl^{-1}$ cross-crosses to make the formula $CaCl_2$ (we don't write 1's as subscripts)
Ex) $Al^{3+} SO_4^{2-}$ makes $Al_2(SO_4)_3$

Write formulas for the following ionic compounds (salts)

- | | |
|-----------------------|-------------------------|
| 1. magnesium fluoride | 6. iron (II) sulfate |
| 2. calcium nitride | 7. barium nitrate |
| 3. aluminum iodide | 8. potassium phosphate |
| 4. sodium chloride | 9. cobalt (II) chloride |
| 5. copper (I) sulfide | 10. zinc carbonate |

Write formulas for the following basic oxides (metal oxides)

- | | |
|--------------------|-----------------------|
| 1. lithium oxide | 6. iron (III) oxide |
| 2. magnesium oxide | 7. silver oxide |
| 3. barium oxide | 8. nickel (II) oxide |
| 4. potassium oxide | 9. tin (IV) oxide |
| 5. cesium oxide | 10. copper (II) oxide |

Write formulas for the following bases (end in hydroxide)

- | | |
|--------------------------|-------------------------|
| 1. lithium hydroxide | 6. lead (II) hydroxide |
| 2. barium hydroxide | 7. zinc hydroxide |
| 3. aluminum hydroxide | 8. calcium hydroxide |
| 4. potassium hydroxide | 9. copper (I) hydroxide |
| 5. copper (II) hydroxide | 10. magnesium hydroxide |

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